



### Naming Compounds & Writing Formulas & Calculating Molar Mass

#### Questions:

1. Identify the following compounds as ionic compound or covalent compound, write the name of the Compounds, and calculate its Molar mass.

	Formula	Ionic or Covalent Compound	Name	Molar Mass (g/mole)
1	NaCl			
2	CO <sub>2</sub>			
3	FeCl <sub>3</sub>			
4	AlF <sub>3</sub>			
5	N <sub>2</sub> O <sub>5</sub>			
6	MgCO <sub>3</sub>			
7	P <sub>2</sub> O <sub>5</sub>			
8	N <sub>2</sub> O <sub>4</sub>			
9	AgNO <sub>3</sub>			
10	Cu <sub>2</sub> O			
11	SO <sub>2</sub>			
12	Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>			
13	IF <sub>5</sub>			
14	P <sub>2</sub> S <sub>3</sub>			
15	Ca(OH) <sub>2</sub>			
16	NaHCO <sub>3</sub>			
17	Na <sub>2</sub> SO <sub>3</sub>			
18	SF <sub>6</sub>			
19	B <sub>2</sub> H <sub>6</sub>			
20	H <sub>2</sub> S			

## Chemistry Worksheet

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2. Identify the following compounds as ionic compound or covalent compound, write the formula of the Compounds, and calculate its Molar mass.

	Name	Ionic or Covalent Compound	Formula	Molar mass (g/mole)
1	Magnesium chloride			
2	Aluminum sulfate			
3	Diboron tetrabromide			
4	Potassium chloride			
5	Sodium fluoride			
6	Sodium chlorite			
7	Sulfur dioxide			
8	Ammonium Chloride			
9	Copper (II) oxide			
10	Nitrogen tribromide			
11	Calcium chloride			
12	Potassium nitrate			
13	Carbone monoxide			
14	Silicon dioxide			
15	Potassium oxide			
16	Tin(IV) selenide			
17	sodium dihydrogen phosphate			
18	Sulfur tetrafluoride			
19	Aluminum oxide			
20	Arsenic pentafluoride			

## Chemistry Worksheet

### Naming Compounds & Writing Formulas & Calculating Molar Mass



### Answers:

1. Identify the following compounds as ionic compound or covalent compound, write the name of the compound, and calculate its molar mass.

	Formula	Ionic or Covalent Compound	Name	Molar Mass (g/mole)
1	NaCl	ionic	sodium chloride	58.45
2	CO <sub>2</sub>	covalent	carbon dioxide	44.01
3	FeCl <sub>3</sub>	ionic	Iron (III) chloride	162.20
4	AlF <sub>3</sub>	ionic	Aluminum fluoride	83.98
5	N <sub>2</sub> O <sub>5</sub>	covalent	dinitrogen pentoxide	108.02
6	MgCO <sub>3</sub>	ionic	Magnesium carbonate	84.32
7	P <sub>2</sub> O <sub>5</sub>	covalent	diphosphorus pentoxide	141.94
8	N <sub>2</sub> O <sub>4</sub>	covalent	dinitrogen tetraoxide	92.02
9	AgNO <sub>3</sub>	ionic	Silver nitrate	169.88
10	Cu <sub>2</sub> O	ionic	Copper (I) oxide	143.10
11	SO <sub>2</sub>	covalent	Sulfur dioxide	64.06
12	Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>	ionic	Calcium phosphate	310.18
13	IF <sub>5</sub>	covalent	Iodine pentafluoride	221.90
14	P <sub>2</sub> S <sub>3</sub>	covalent	diphosphorous trisulfide	158.12
15	Ca(OH) <sub>2</sub>	ionic	Calcium hydroxide	74.10
16	NaHCO <sub>3</sub>	ionic	Sodium hydrogen carbonate	84.01
17	Na <sub>2</sub> SO <sub>3</sub>	ionic	Sodium sulfite	126.04
18	SF <sub>6</sub>	covalent	Sulfur hexafluoride	146.06
19	B <sub>2</sub> H <sub>6</sub>	covalent	diboron hexahydride	27.68
20	H <sub>2</sub> S	covalent	dihydrogen monosulfide	34.08

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2. Identify the following compounds as ionic compound or covalent compound, write the formula of the compound, and Calculate its Molar mass.

	Name	Ionic or Covalent Compound	Formula	Molar mass (g/mole)
1	Magnesium chloride	Ionic	MgCl <sub>2</sub>	95.21
2	Aluminum sulfate	Ionic	Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	342.14
3	Diboron tetrabromide	ionic	B <sub>2</sub> Br <sub>4</sub>	341.22
4	Potassium chloride	Ionic	KCl	74.55
5	Sodium fluoride	Ionic	NaF	41.99
6	Sodium chlorite	Ionic	NaClO <sub>2</sub>	90.44
7	Sulfur dioxide	covalent	SO <sub>2</sub>	64.06
8	Ammonium Chloride	Ionic	NH <sub>4</sub> Cl	53.50
9	Copper (II) oxide	Ionic	CuO	79.55
10	Nitrogen tribromide	covalent	NBr <sub>3</sub>	253.71
11	Calcium chloride	Ionic	CaCl <sub>2</sub>	110.98
12	Potassium nitrate	Ionic	KNO <sub>3</sub>	101.11
13	Carbone monoxide	covalent	CO	28.01
14	Silicon dioxide	covalent	SiO <sub>2</sub>	60.09
15	Potassium oxide	Ionic	KO <sub>2</sub>	71.10
16	Tin(IV) selenide	Ionic	SnSe <sub>2</sub>	276.63
17	sodium dihydrogen phosphate	Ionic	NaH <sub>2</sub> PO <sub>4</sub>	119.98
18	Sulfur tetrafluoride	covalent	SF <sub>4</sub>	108.06
19	Aluminum oxide	Ionic	Al <sub>2</sub> O <sub>3</sub>	101.96
20	Arsenic pentfluoride	covalent	AsF <sub>5</sub>	169.91