

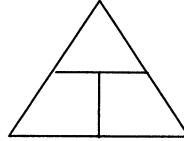
## Density Practice: Worksheet #1

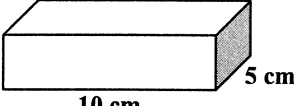
**Calculate density, and identify substances using a density chart.**

Density is a measure of the amount of mass in a certain volume. This physical property is often used to identify and classify substances. It is usually expressed in grams per cubic centimeters, or  $\text{g/cm}^3$ . The chart on the right lists the densities of some common materials.

Substance	Density ( $\text{g/cm}^3$ )
Gold	19.3
Mercury	13.5
Lead	11.4
Iron	7.87
Aluminum	3.7
Bone	1.7-2.0
Gasoline	0.66-0.69
Air (dry)	0.00119

Equation: **Density =  $\frac{\text{mass}}{\text{Volume}}$  or  $D = \frac{m}{V}$**



Problem Statement	Formula	Define Variables	Substitution	Answer
<i>Sample:</i> What is the density of a billiard ball that has a volume of $100 \text{ cm}^3$ and a mass of 250 g?	$D = \frac{m}{V}$	$M = 250 \text{ g}$ $V = 100 \text{ cm}^3$	$D = \frac{250 \text{ g}}{100 \text{ cm}^3}$	$2.5 \text{ g/cm}^3$
1. A loaf of bread has a volume of $2270 \text{ cm}^3$ and a mass of 454 g. What is the density of the bread?	$D = \frac{m}{V}$	$V = 2270 \text{ cm}^3$ $m = 454 \text{ g}$	$D = \frac{454 \text{ g}}{2270 \text{ cm}^3}$	$.2 \text{ g/cm}^3$
2. A block of wood has a density of $0.6 \text{ g/cm}^3$ and a volume of $1.2 \text{ cm}^3$ . What is the mass of the block of wood?	$D = \frac{m}{V}$	$D = .6 \text{ g/cm}^3$ $V = 1.2 \text{ cm}^3$	$0.6 = \frac{m}{1.2 \text{ cm}^3}$ $m = (.6)(1.2)$	$.72 \text{ g}$
3. A 800g boulder has a density of $8 \text{ g/cm}^3$ . What is the volume of the boulder?	$D = \frac{m}{V}$	$m = 800 \text{ g}$ $D = 8 \text{ g/cm}^3$	$8 = \frac{800}{V}$ $V = \frac{800}{8}$	$100 \text{ cm}^3$
4. What is the mass of the block of iron illustrated below? 	$D = \frac{m}{V}$ $m = (D)(V)$	$V = 100 \text{ cm}^3$ $D = 7.87 \text{ g/cm}^3$	$(100 \text{ cm}^3)(7.87 \text{ g/cm}^3)$	$m = 287 \text{ g}$

Use the data below to calculate the density of each unknown substance. Then use the density chart above to determine the identity of each substance.

Mass (g)	Volume ( $\text{cm}^3$ )	$D = m/v$ Variable Substitutions	Density ( $\text{g/cm}^3$ )	Substance
4725	350	$D = \frac{4725}{350}$	$D = 13.5$	Mercury
171	15	$D = \frac{171}{15}$	11.4	Lead
148	40	$D = \frac{148}{40}$	3.7	Aluminum
475	250	$D = \frac{475}{250}$	1.9	Bone
680	1000	$D = \frac{680}{1000}$	.68	Gasoline